**Summary of Different Types of Carbon−Carbon Covalent Bonds**





Other Key Points:

* The shorter the bond, the stronger the bond
* The greater the electron density in the region of orbital overlap, the stronger the bond
* The more s character in the orbital used for overlap, the shorter and the stronger the bond
* The more s character in the orbital used for overlap, the larger the bond angle
* The more s character in the orbital, the greater the electronegativity of that atom